

<b>hydrogen ion</b>	<b>lithium ion</b>	<b>sodium ion</b>
<b>potassium ion</b>	<b>rubidium ion</b>	<b>cesium ion</b>
<b>beryllium ion</b>	<b>magnesium ion</b>	<b>calcium ion</b>
<b>strontium ion</b>	<b>barium ion</b>	<b>radium ion</b>
<b>aluminum ion</b>	<b>hydride ion</b>	<b>fluoride ion</b>
<b>chloride ion</b>	<b>bromide ion</b>	<b>iodide ion</b>
<b>oxide ion</b>	<b>sulfide ion</b>	<b>nitride ion</b>

<b>Na<sup>+</sup></b>	<b>Li<sup>+</sup></b>	<b>H<sup>+</sup></b>
<b>Cs<sup>+</sup></b>	<b>Rb<sup>+</sup></b>	<b>K<sup>+</sup></b>
<b>Ca<sup>2+</sup></b>	<b>Mg<sup>2+</sup></b>	<b>Be<sup>2+</sup></b>
<b>Ra<sup>2+</sup></b>	<b>Ba<sup>2+</sup></b>	<b>Sr<sup>2+</sup></b>
<b>F<sup>-</sup></b>	<b>H<sup>-</sup></b>	<b>Al<sup>3+</sup></b>
<b>I<sup>-</sup></b>	<b>Br<sup>-</sup></b>	<b>Cl<sup>-</sup></b>
<b>N<sup>3-</sup></b>	<b>S<sup>2-</sup></b>	<b>O<sup>2-</sup></b>

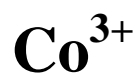
<b>phosphide ion</b>	<b>nickel ion</b> Roman numeral not needed	<b>zinc ion</b> Roman numeral not needed
<b>silver ion</b> Roman numeral not needed	<b>hydronium ion</b>	<b>ammonium ion</b>
<b>Roman numeral for 3</b>	<b>Roman numeral for 2</b>	<b>Roman numeral for 1</b>
<b>Roman numeral for 6</b>	<b>Roman numeral for 5</b>	<b>Roman numeral for 4</b>
<b>Roman numeral for 8</b>	<b>Roman numeral for 7</b>	<b>titanium(IV) ion</b> example of Roman numeral usage
<b>chromium(II) ion</b> <b>chromous ion</b>	<b>chromium(III) ion</b> <b>chromic ion</b>	<b>manganese(II) ion</b>
<b>manganese(III) ion</b>	<b>iron(II) ion</b> <b>ferrous ion</b>	<b>iron(III) ion</b> <b>ferric ion</b>

<p style="text-align: center;"><b>Zn<sup>2+</sup></b></p> <p style="text-align: center;">Roman numeral not needed</p>	<p style="text-align: center;"><b>Ni<sup>2+</sup></b></p> <p style="text-align: center;">Roman numeral not needed</p>	<p style="text-align: center;"><b>P<sup>3-</sup></b></p>
<p style="text-align: center;"><b>NH<sub>4</sub><sup>+</sup></b></p>	<p style="text-align: center;"><b>H<sub>3</sub>O<sup>+</sup></b></p>	<p style="text-align: center;"><b>Ag<sup>+</sup></b></p> <p style="text-align: center;">Roman numeral not needed</p>
<p style="text-align: center;"><b>I</b></p>	<p style="text-align: center;"><b>II</b></p>	<p style="text-align: center;"><b>III</b></p>
<p style="text-align: center;"><b>IV</b></p>	<p style="text-align: center;"><b>V</b></p>	<p style="text-align: center;"><b>VI</b></p>
<p style="text-align: center;"><b>Ti<sup>4+</sup></b></p> <p style="text-align: center;">example of Roman numeral usage</p>	<p style="text-align: center;"><b>VII</b></p>	<p style="text-align: center;"><b>VIII</b></p>
<p style="text-align: center;"><b>Mn<sup>2+</sup></b></p>	<p style="text-align: center;"><b>Cr<sup>3+</sup></b></p> <p style="text-align: center;">two names</p>	<p style="text-align: center;"><b>Cr<sup>2+</sup></b></p> <p style="text-align: center;">two names</p>
<p style="text-align: center;"><b>Fe<sup>3+</sup></b></p> <p style="text-align: center;">two names</p>	<p style="text-align: center;"><b>Fe<sup>2+</sup></b></p> <p style="text-align: center;">two names</p>	<p style="text-align: center;"><b>Mn<sup>3+</sup></b></p>

<b>cobalt(II) ion</b> <b>cobaltous ion</b>	<b>cobalt(III) ion</b> <b>cobaltic ion</b>	<b>copper(I) ion</b> <b>cuprous ion</b>
<b>copper(II) ion</b> <b>cupric ion</b>	<b>tin(II) ion</b> <b>stannous ion</b>	<b>tin(IV) ion</b> <b>stannic ion</b>
<b>mercury(I) ion</b> <b>mercurous ion</b>	<b>mercury(II) ion</b> <b>mercuric ion</b>	<b>gold(I) ion</b>
<b>gold(III) ion</b>	<b>lead(II) ion</b> <b>plumbous ion</b>	<b>lead(IV) ion</b> <b>plumbic ion</b>
<b>hypochlorite ion</b>	<b>hypobromite ion</b>	<b>hypoiodite ion</b>
<b>chlorite ion</b>	<b>bromite ion</b>	<b>iodite ion</b>
<b>chlorate ion</b>	<b>bromate ion</b>	<b>iodate ion</b>



two names



two names



two names



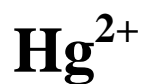
two names



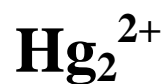
two names



two names



two names



two names



two names



two names



<b>perchlorate ion</b>	<b>perbromate ion</b>	<b>periodate ion</b>
<b>hydroxide ion</b>	<b>cyanide ion</b>	<b>peroxide ion</b>
<b>thiocyanate ion</b>	<b>oxalate ion</b>	<b>permanganate ion</b>
<b>hydrogen carbonate ion</b> <b>bicarbonate ion</b>	<b>carbonate ion</b>	<b>acetate ion</b>
<b>dihydrogen phosphate ion</b>	<b>hydrogen phosphate ion</b>	<b>phosphate ion</b>
<b>sulfite ion</b>	<b>hydrogen sulfite ion</b> <b>bisulfite ion</b>	<b>sulfate ion</b>
<b>hydrogen sulfate ion</b> <b>bisulfate ion</b>	<b>nitrite ion</b>	<b>nitrate ion</b>

$\text{IO}_4^-$	$\text{BrO}_4^-$	$\text{ClO}_4^-$
$\text{O}_2^{2-}$	$\text{CN}^-$	$\text{OH}^-$
$\text{MnO}_4^-$	$\text{C}_2\text{O}_4^{2-}$	$\text{SCN}^-$
$\text{CH}_3\text{COO}^-$ $\text{CH}_3\text{CO}_2^-$ $\text{C}_2\text{H}_3\text{O}_2^-$	$\text{CO}_3^{2-}$	$\text{HCO}_3^-$ two names
$\text{PO}_4^{3-}$	$\text{HPO}_4^{2-}$	$\text{H}_2\text{PO}_4^-$
$\text{SO}_4^{2-}$	$\text{HSO}_3^-$ two names	$\text{SO}_3^{2-}$
$\text{NO}_3^-$	$\text{NO}_2^-$	$\text{HSO}_4^-$ two names



<b>chromate ion</b>	<b>dichromate ion</b>	<b>prefix for 1</b>
<b>prefix for 2</b>	<b>prefix for 3</b>	<b>prefix for 4</b>
<b>prefix for 5</b>	<b>prefix for 6</b>	<b>prefix for 7</b>
<b>prefix for 8</b>	<b>prefix for 9</b>	<b>prefix for 10</b>
<b>tetraphosphorus heptasulfide</b> binary nonmetal example	<b>nitrogen triiodide</b> binary nonmetal example	<b>dinitrogen pentoxide</b> binary nonmetal example
<b>hydrochloric acid</b>	<b>hydrobromic acid</b>	<b>hydroiodic acid</b>
<b>hypochlorous acid</b>	<b>chlorous acid</b>	<b>chloric acid</b>

<b>mono-</b>	<b>Cr<sub>2</sub>O<sub>7</sub><sup>2-</sup></b>	<b>CrO<sub>4</sub><sup>2-</sup></b>
<b>tetra-</b>	<b>tri-</b>	<b>di-</b>
<b>hepta-</b>	<b>hexa-</b>	<b>penta-</b>
<b>deca-</b>	<b>nona-</b>	<b>octa-</b>
<b>N<sub>2</sub>O<sub>5</sub></b> contract 'pentaoxide'	<b>NI<sub>3</sub></b> no mono on the first element, do not contract 'triiodide'	<b>P<sub>4</sub>S<sub>7</sub></b>
<b>HI(aq)</b>	<b>HBr(aq)</b>	<b>HCl(aq)</b>
<b>HClO<sub>3</sub>(aq)</b>	<b>HClO<sub>2</sub>(aq)</b>	<b>HClO(aq)</b>

<b>hypobromous acid</b>	<b>bromous acid</b>	<b>bromic acid</b>
<b>hypoiodous acid</b>	<b>iodous acid</b>	<b>iodic acid</b>
<b>perbromic acid</b>	<b>periodic acid</b>	<b>perchloric acid</b>
<b>nitrous acid</b>	<b>sulfurous acid</b>	<b>phosphorous acid</b>
<b>nitric acid</b>	<b>sulfuric acid</b>	<b>phosphoric acid</b>
<b>oxalic acid</b>	<b>hydrocyanic acid</b>	<b>carbonic acid</b>
<b>hydrosulfuric acid</b>	<b>acetic acid</b>	<b>hydrofluoric acid</b>

<b><math>\text{HBrO}_3(aq)</math></b>	<b><math>\text{HBrO}_2(aq)</math></b>	<b><math>\text{HBrO}(aq)</math></b>
<b><math>\text{HIO}_3(aq)</math></b>	<b><math>\text{HIO}_2(aq)</math></b>	<b><math>\text{HIO}(aq)</math></b>
<b><math>\text{HClO}_4(aq)</math></b>	<b><math>\text{HIO}_4(aq)</math></b>	<b><math>\text{HBrO}_4(aq)</math></b>
<b><math>\text{H}_3\text{PO}_3(aq)</math></b>	<b><math>\text{H}_2\text{SO}_3(aq)</math></b>	<b><math>\text{HNO}_2(aq)</math></b>
<b><math>\text{H}_3\text{PO}_4(aq)</math></b>	<b><math>\text{H}_2\text{SO}_4(aq)</math></b>	<b><math>\text{HNO}_3(aq)</math></b>
<b><math>\text{H}_2\text{CO}_3(aq)</math></b>	<b><math>\text{HCN}(aq)</math></b>	<b><math>\text{H}_2\text{C}_2\text{O}_4(aq)</math></b>
<b><math>\text{HF}(aq)</math></b>	<b><math>\text{CH}_3\text{COOH}(aq)</math> <b><math>\text{HCH}_3\text{CO}_2(aq)</math> <b><math>\text{HC}_2\text{H}_3\text{O}_2(aq)</math></b></b></b>	<b><math>\text{H}_2\text{S}(aq)</math></b>

<p><b>ammonia</b></p> <p>nonsystematic</p>	<p><b>water</b></p> <p>nonsystematic</p>	<p><b>methane</b></p> <p>nonsystematic</p>
<p><b>copper(II) sulfate pentahydrate</b></p> <p>example of a hydrated salt</p>	<p><b>ozone</b></p> <p>nonsystematic</p>	<p><b>hydrogen</b></p> <p><b>helium</b></p> <p><b>lithium</b></p>
<p><b>beryllium</b></p> <p><b>boron</b></p> <p><b>carbon</b></p>	<p><b>nitrogen</b></p> <p><b>oxygen</b></p> <p><b>fluorine</b></p>	<p><b>neon</b></p> <p><b>sodium</b></p> <p><b>magnesium</b></p>
<p><b>aluminum</b></p> <p><b>silicon</b></p> <p><b>phosphorus</b></p>	<p><b>sulfur</b></p> <p><b>chlorine</b></p> <p><b>argon</b></p>	<p><b>potassium</b></p> <p><b>calcium</b></p> <p><b>titanium</b></p>
<p><b>chromium</b></p> <p><b>manganese</b></p> <p><b>iron</b></p>	<p><b>cobalt</b></p> <p><b>nickel</b></p> <p><b>copper</b></p>	<p><b>zinc</b></p> <p><b>arsenic</b></p> <p><b>selenium</b></p>
<p><b>bromine</b></p> <p><b>krypton</b></p> <p><b>rubidium</b></p>	<p><b>strontium</b></p> <p><b>molybdenum</b></p> <p><b>silver</b></p>	<p><b>cadmium</b></p> <p><b>tellurium</b></p> <p><b>iodine</b></p>
<p><b>cesium</b></p> <p><b>barium</b></p> <p><b>lanthanum</b></p>	<p><b>tungsten</b>    <b>gold</b></p> <p><b>mercury</b></p> <p><b>radium</b></p>	<p><b>tin</b>            <b>lead</b></p> <p><b>actinium</b></p> <p><b>uranium</b></p>

<b>CH<sub>4</sub></b> nonsystematic	<b>H<sub>2</sub>O</b> nonsystematic	<b>NH<sub>3</sub></b> nonsystematic
<b>H</b> <b>He</b> <b>Li</b>	<b>O<sub>3</sub></b> nonsystematic	<b>CuSO<sub>4</sub>·5H<sub>2</sub>O</b> example of a hydrated salt
<b>Ne</b> <b>Na</b> <b>Mg</b>	<b>N</b> <b>O</b> <b>F</b>	<b>Be</b> <b>B</b> <b>C</b>
<b>K</b> <b>Ca</b> <b>Ti</b>	<b>S</b> <b>Cl</b> <b>Ar</b>	<b>Al</b> <b>Si</b> <b>P</b>
<b>Zn</b> <b>As</b> <b>Se</b>	<b>Co</b> <b>Ni</b> <b>Cu</b>	<b>Cr</b> <b>Mn</b> <b>Fe</b>
<b>Cd</b> <b>Te</b> <b>I</b>	<b>Sr</b> <b>Mo</b> <b>Ag</b>	<b>Br</b> <b>Kr</b> <b>Rb</b>
<b>Sn</b> <b>Pb</b>	<b>W</b> <b>Au</b>	<b>Cs</b>
<b>Ac</b> <b>U</b>	<b>Hg</b> <b>Ra</b>	<b>Ba</b> <b>La</b>