

### SI Units: Base (metric) [derived]

Unit of...	Name	Symbol
Length	meter	m
Mass	kilogram	kg
	(gram)	(g)
Time	second	s
Temperature	kelvin	K
	(Celsius)	(°C)
Amount	mole	mol
[Energy]	[Joule]	[J]

### SI Power of 10 Prefixes

Factor	Name	Symbol	Factor	Name	Symbol
			$10^{-1}$	deci	d
$10^9$	giga	G	$10^{-2}$	centi	c
$10^6$	mega	M	$10^{-3}$	milli	m
$10^3$	kilo	k	$10^{-6}$	micro	$\mu$
			$10^{-9}$	nano	n

### Numerical Prefixes

- 1 = mono
- 2 = di
- 3 = tri
- 4 = tetra
- 5 = penta
- 6 = hexa
- 7 = hepta
- 8 = octa
- 9 = nona
- 10 = deca

Exact equivalences:

$1 \text{ mL} = 1 \text{ cm}^3$

$1 \text{ in} = 2.54 \text{ cm}$

$1 \text{ cal} = 4.184 \text{ J}$

$10 = \text{deca}$

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<td>oxygen</td> <td>fluorine</td> <td>neon</td> </tr> <tr> <td>3</td> <td colspan="2">11</td> <td colspan="2">12</td> <td colspan="2">3</td> <td colspan="2">4</td> <td colspan="2">5</td> <td colspan="2">6</td> <td colspan="2">7</td> <td colspan="2">8</td> <td colspan="2">9</td> <td colspan="2">10</td> <td colspan="2">11</td> <td colspan="2">12</td> <td colspan="2">13</td> <td colspan="2">14</td> <td colspan="2">15</td> <td colspan="2">16</td> <td colspan="2">17</td> <td colspan="2">18</td> </tr> <tr> <td><b>Na</b></td> <td><b>Mg</b></td> <td colspan="16"></td> <td><b>Al</b></td> <td><b>Si</b></td> <td><b>P</b></td> <td><b>S</b></td> <td><b>Cl</b></td> <td><b>Ar</b></td> </tr> <tr> <td>sodium</td> <td>magnesium</td> <td colspan="16"></td> <td>aluminum</td> <td>silicon</td> <td>phosphorus</td> <td>sulfur</td> <td>chlorine</td> <td>argon</td> </tr> <tr> <td>4</td> <td colspan="2">19</td> <td colspan="2">20</td> <td colspan="2">21</td> <td colspan="2">22</td> <td colspan="2">23</td> <td 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<td>gold</td> <td>mercury</td> <td>thallium</td> <td>lead</td> <td>bismuth</td> <td>polonium</td> <td>astatine</td> <td>radon</td> </tr> <tr> <td>7</td> <td colspan="2">87</td> <td colspan="2">88</td> <td colspan="2">89</td> <td colspan="2">104</td> <td colspan="2">105</td> <td colspan="2">106</td> <td colspan="2">107</td> <td colspan="2">108</td> <td colspan="2">109</td> <td colspan="2">110</td> <td colspan="2">111</td> <td colspan="2">112</td> <td colspan="2">113</td> <td colspan="2">114</td> <td colspan="2">115</td> <td colspan="2">116</td> <td colspan="2">117</td> <td colspan="2">118</td> </tr> <tr> <td><b>Fr</b></td> <td><b>Ra</b></td> <td><b>Ac</b></td> <td><b>Rf</b></td> <td><b>Db</b></td> <td><b>Sg</b></td> <td><b>Bh</b></td> <td><b>Hs</b></td> <td><b>Mt</b></td> <td><b>Ds</b></td> <td><b>Rg</b></td> <td><b>Cn</b></td> <td><b>Nh</b></td> <td><b>Fl</b></td> <td><b>Mc</b></td> <td><b>Lv</b></td> <td><b>Ts</b></td> <td><b>Og</b></td> </tr> <tr> <td>francium</td> <td>radium</td> 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### Monatomic Nonmetal Anions

Description	Example	Name
Group 17 = 1-	Cl <sup>-</sup>	chloride
Group 16 = 2-	S <sup>2-</sup>	sulfide
Group 15 = 3-	N <sup>3-</sup>	nitride

### Monatomic Metal Cations

Description	Example	Name
Group 1 = 1+	Na <sup>+</sup>	sodium
Group 2 = 2+	Ca <sup>2+</sup>	calcium
Group 13 = 3+	Al <sup>3+</sup>	aluminum
Transition = Roman numeral	Au <sup>3+</sup>	gold(III)
Inner transition = Roman num.	U <sup>6+</sup>	uranium(VI)

### Alternative (old) names for some transition metal ions

Cr <sup>2+</sup> = chromous	Cr <sup>3+</sup> = chromic
Fe <sup>2+</sup> = ferrous	Fe <sup>3+</sup> = ferric
Co <sup>2+</sup> = cobaltous	Co <sup>3+</sup> = cobaltic
Cu <sup>+</sup> = cuprous	Cu <sup>2+</sup> = cupric
Sn <sup>2+</sup> = stannous	Sn <sup>4+</sup> = stannic
Hg <sub>2</sub> <sup>2+</sup> = mercurous	Hg <sup>2+</sup> = mercuric
Pb <sup>2+</sup> = plumbous	Pb <sup>4+</sup> = plumbic

### Naming Binary Nonmetal Compounds:

First element followed by ionic name of second (use prefixes).  
no *mono* on first element; shorten double vowels *ao* → *o*; *oo* → *o*

### Polyatomic Ions

Formula	Name	Formula	Name
NH <sub>4</sub> <sup>+</sup>	ammonium	CO <sub>3</sub> <sup>2-</sup>	carbonate
H <sub>3</sub> O <sup>+</sup>	hydronium	HCO <sub>3</sub> <sup>-</sup>	hydrogen carbonate or bicarbonate
CN <sup>-</sup>	cyanide	PO <sub>3</sub> <sup>3-</sup>	phosphite
SCN <sup>-</sup>	thiocyanate	PO <sub>4</sub> <sup>3-</sup>	phosphate
OH <sup>-</sup>	hydroxide	HPO <sub>4</sub> <sup>2-</sup>	hydrogen phosphate
O <sub>2</sub> <sup>2-</sup>	peroxide	H <sub>2</sub> PO <sub>4</sub> <sup>-</sup>	dihydrogen phosphate
CH <sub>3</sub> COO <sup>-</sup>	acetate	ClO <sup>-</sup>	hypochlorite
CH <sub>3</sub> CO <sub>2</sub> <sup>-</sup>		ClO <sub>2</sub> <sup>-</sup>	chlorite
C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> <sup>-</sup>	chromate	ClO <sub>3</sub> <sup>-</sup>	chlorate
CrO <sub>4</sub> <sup>2-</sup>	dichromate	ClO <sub>4</sub> <sup>-</sup>	perchlorate
Cr <sub>2</sub> O <sub>7</sub> <sup>2-</sup>	permanganate	BrO <sup>-</sup>	hypobromite
MnO <sub>4</sub> <sup>-</sup>	nitrite	BrO <sub>2</sub> <sup>-</sup>	bromite
NO <sub>2</sub> <sup>-</sup>	nitrate	BrO <sub>3</sub> <sup>-</sup>	bromate
NO <sub>3</sub> <sup>-</sup>	sulfite	BrO <sub>4</sub> <sup>-</sup>	perbromate
SO <sub>3</sub> <sup>2-</sup>	sulfate	IO <sup>-</sup>	hypoiodite
SO <sub>4</sub> <sup>2-</sup>	hydrogen sulfate or bisulfate	IO <sub>2</sub> <sup>-</sup>	iodite
HSO <sub>4</sub> <sup>-</sup>	oxalate	IO <sub>3</sub> <sup>-</sup>	iodate
C <sub>2</sub> O <sub>4</sub> <sup>2-</sup>		IO <sub>4</sub> <sup>-</sup>	periodate

### Naming Acids from Anions:

- ide → hydro<root>ic acid
  - ite → <root>ous acid
  - ate → <root>ic acid
- bromide HBr = hydrobromic acid  
chlorite HClO<sub>2</sub> = chlorous acid  
nitrate HNO<sub>3</sub> = nitric acid

